

LMS Use During Academic Year 2020-21

- Moodle

The screenshot shows a web browser window with the URL `sutarmath.gnomio.com`. The browser's address bar and tabs are visible at the top. The Moodle interface includes a navigation menu on the left with the site name "Moodle". A top right notification states "You are not logged in. (Log in)". The main content area features a large header with the site name, a welcome message, and a list of links of interest. Below this, a section titled "Available courses" lists several courses, each with a course icon, name, and a right-pointing arrow. The courses listed are: TYBSC 2021-22- (MTH-502: Real Analysis-I), TYBSC 2021-22- (MTH-503: Algebra), MSC-I (2021-22), TYBSC 2020-21, SYBSC-GROUP THEORY[MTH-302(A)], and MTH-366(A): Applied Numerical Methods. The last course entry also includes the text "Teacher: Dr. Chandrashekhhar Sutar".

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- [Moodle manuals](#)
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Available courses

- TYBSC 2021-22- (MTH-502: Real Analysis-I)
- TYBSC 2021-22- (MTH-503: Algebra)
- MSC-I (2021-22)
- TYBSC 2020-21
- SYBSC-GROUP THEORY[MTH-302(A)]
- MTH-366(A): Applied Numerical Methods

Teacher: Dr. Chandrashekhhar Sutar

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Dr. Milind Patil

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+74 in last 28 days

Summary
Last 28 days

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Top videos
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TYBSc CH-504 U-II L-2	137
TYBSc CH-504 U-II L-3	107
PG Pathshala CH-350 ORM Introduction Lecture L-1	60

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FYBSc CH-101 Physical and Inorganic Chemistry-I (SEM-I)...	FYBSc CH-102 Organic and Inorganic Chemistry-I (SEM-I)...	SYBSc CH-301 Physical and Inorganic Chemistry (SEM-II)...	SYBSc CH-302 Organic and Inorganic Chemistry (SEM-II)...	SYBSc CH-304 Basic Analytical Chemistry (SEC-1) (SEM-II)...
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TYBSc CH-501 Principles of Physical Chemistry-I (SEM-V)...	TYBSc CH-502 Inorganic Chemistry (SEM-V) (KBCNMU)	TYBSc CH-503 Organic Reaction Mechanism (SEM-V)...	TYBSc CH-504 Industrial Chemistry (Sem-V) (KBCNMU)	TYBSc CH-505 Analytical Instrumentation (SEC)...
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GREEN CHEMISTRY 15	TUTORIAL : ZOOM APP 5
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- Google Classroom

Google Classroom



To-do To review Calendar

CH-O-4 Short Resea...
2020-21

CH-O-2 Ternary Mixt...
2020-21

TYBSc (Green Chemi...
Sem-V (Academic Year 2021-22)

TYBSc (Industrial Ch...
Sem-V (Academic Year 2021-22)

MSc Part-I (Name Re...
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Sem-VI (Academic Year 2020-21)

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MSc Part-II (Heterocy...
Sem-IV (Academic Year 2020-21)

Faculty Development...
EDP Vigyan Ashram

MSc Part-II (Organic ...
Sem-III (Academic Year 2020-21)

MSc Part-II (Hetero C...
Sem-IV (Academic Year 2019-20)

PSGVPM's ASC Colle...
Academic and Administration
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MSc Part-II (Free Rad...
Sem-III (Academic Year 2020-21)

TYBSc (Green Chemi...
Sem-V (Academic Year 2020-21)

TYBSc (Industrial Ch...
Sem-V (Academic Year 2020-21)

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Broadcasts

Destinations

Broadcasts

Upcoming

Past

Title	Started	Duration	
Diels Alder Reaction Organic Name R... Ended	Jul 16, 2021, 10:29 AM	33 min 53 sec	⋮
Grignard reaction Organic Name Rea... Ended	Jul 16, 2021, 09:14 AM	33 min 39 sec	⋮
Higher Studies & Career Prospects in Sc... Ended	Jul 10, 2021, 05:03 PM	3 hour 11 min 14 sec	⋮
Reformatsky reaction with solution Tric... Ended	Jul 10, 2021, 10:41 AM	21 min 25 sec	⋮
Reimer-Tiemann reaction Organic N... Ended	Jul 9, 2021, 05:07 PM	20 min 53 sec	⋮
Dieckmann condensation Ended	Jul 8, 2021, 08:18 PM	15 min 23 sec	⋮

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Destinations

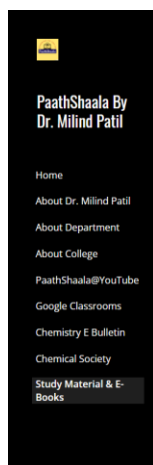
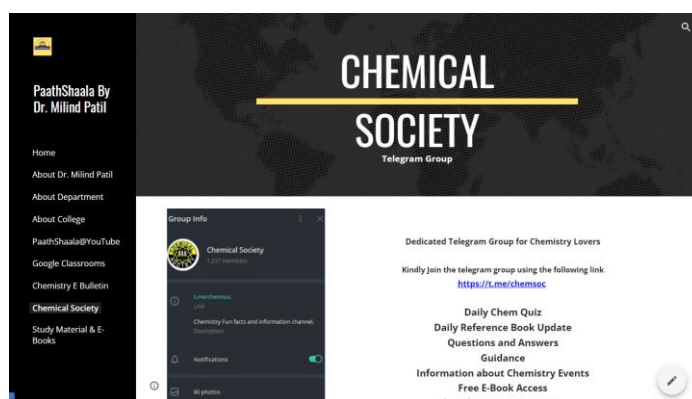
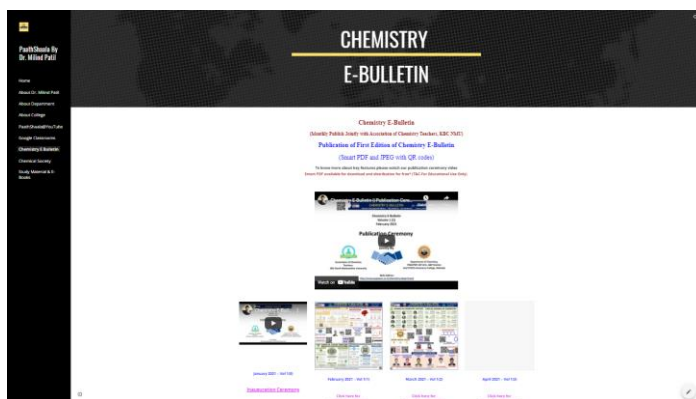
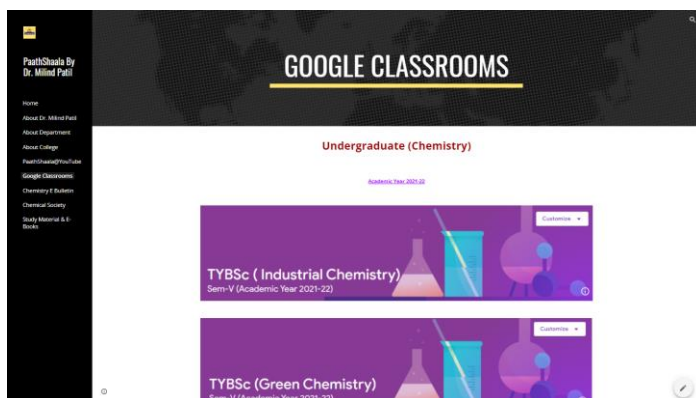
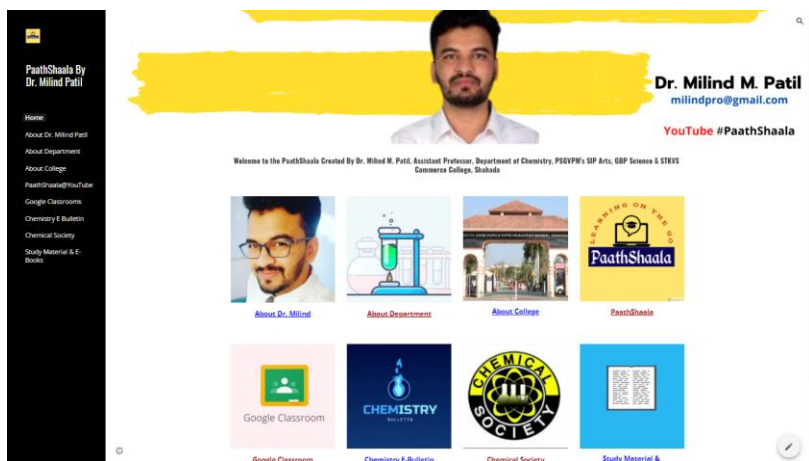
Higher Studies & Career Prospects in Sc... Ended	Jul 10, 2021, 05:03 PM	3 hour 11 min 14 sec	⋮
Reformatsky reaction with solution Tric... Ended	Jul 10, 2021, 10:41 AM	21 min 25 sec	⋮
Reimer-Tiemann reaction Organic N... Ended	Jul 9, 2021, 05:07 PM	20 min 53 sec	⋮
Dieckmann condensation Ended	Jul 8, 2021, 08:18 PM	15 min 23 sec	⋮
Stobbe Condensation Organic Name ... Ended	Jul 7, 2021, 07:01 PM	32 min 14 sec	⋮
Bayer-Villiger Oxidation Organic Na... Ended	Jul 2, 2021, 02:11 PM	1 hour 3 min 14 sec	⋮
Introduction of CH-250 Ended	Jul 1, 2021, 12:10 PM	44 min 28 sec	⋮
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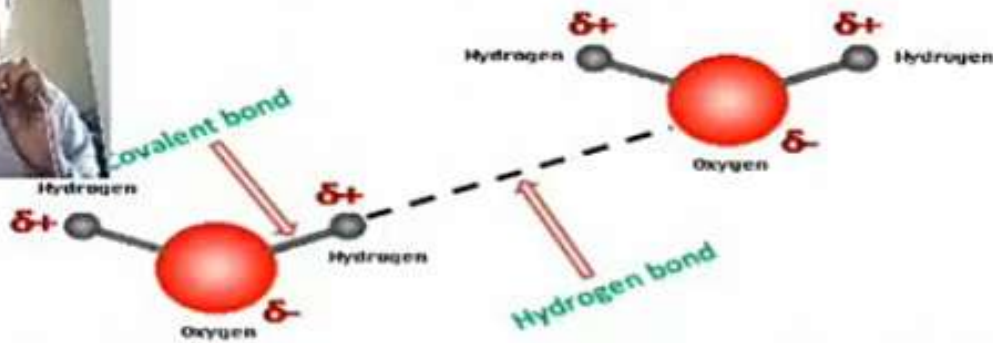
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TITLE	LAST MODIFIED
Assignment	7/13/19 MadChem Classes Chemistry
PDF Notes	6/14/19 MadChem Classes Chemistry
(CSIR&GATE)Assignment Atom structure & Quantum chemist...	2/6/19 MadChem Classes Chemistry
(JAM)Assignment inorganic chemistry 1.docx	3/25/19 MadChem Classes Chemistry
(NET-GATE)Assignment inorganic chemistry 1 (2).pdf	2/6/19 MadChem Classes Chemistry
Assignment Chemical Kinetics for CSIR-NET.docx	3/25/19 MadChem Classes Chemistry
Assignment organometallic compounds.pdf	1/30/19 MadChem Classes Chemistry

Chemistry Competitive Examination Study Material (NET/SET/GATE/IIT-JAM etc.)

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Important Topic	12/14/20 Amit Gavil

Hydrogen Bond



The interaction between the oppositely charged ends of permanently polarized molecules each containing a hydrogen atom is called as **hydrogen bond**.

Characteristics:

- Cohesive energy: 0.5 eV per bond.
- Materials exhibiting hydrogen bond possess high M.P and B.P compared with molecular solids.



Description



PHY 502 Unit 3 Lecture 4



Prashant Jagtap

1

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Calculation of Ripple factor

• **Ripple factor** is defined as the ratio of root mean square (r.m.s) value of a.c. component to the value of d.c. Component.

$$\therefore \text{Ripple factor} = \frac{\text{rms value of a.c. component}}{\text{value of d.c component}}$$

$$\therefore \text{Ripple factor} = \frac{I_{a.c.}}{I_{d.c.}} \quad \text{————— (1)}$$

By definition, the r.m.s value of load current is given by

$$I_{rms} = \sqrt{I_{a.c}^2 + I_{d.c}^2}$$

$$I_{rms}^2 = I_{a.c}^2 + I_{d.c}^2$$

$$I_{a.c}^2 = I_{rms}^2 - I_{d.c}^2$$

$$\therefore I_{a.c} = \sqrt{I_{rms}^2 - I_{d.c}^2}$$



Description



PHY 302A Unit 2 Lecture3



Prashant Jagtap

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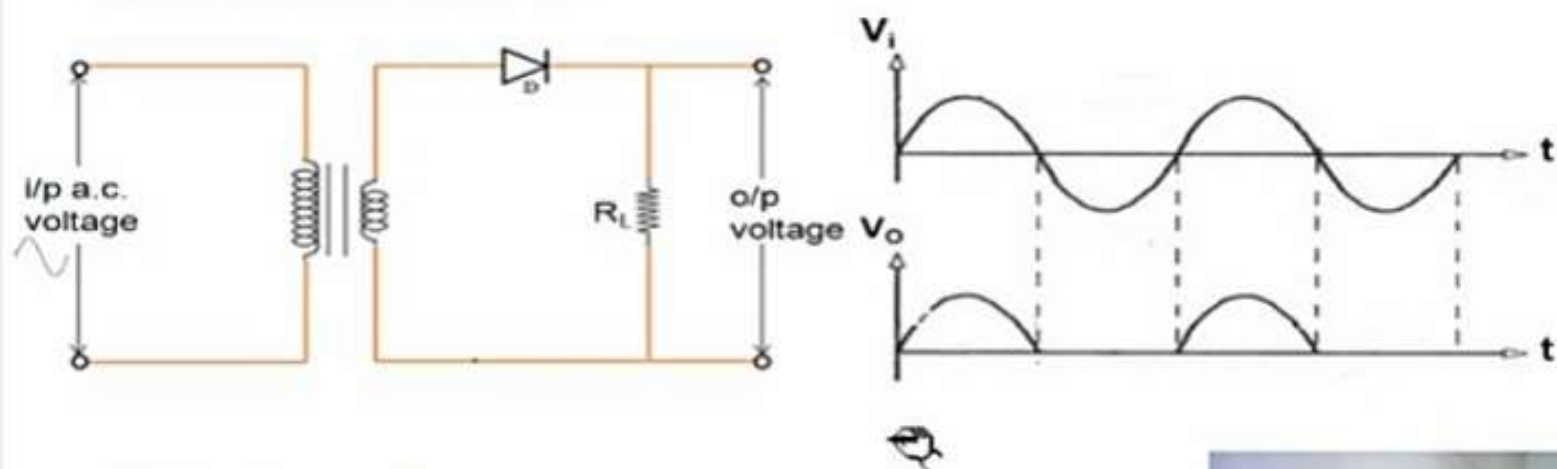
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1) Half wave rectifier:



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PHY 302A Unit 2 Lecture 1



Prashant Jagtap

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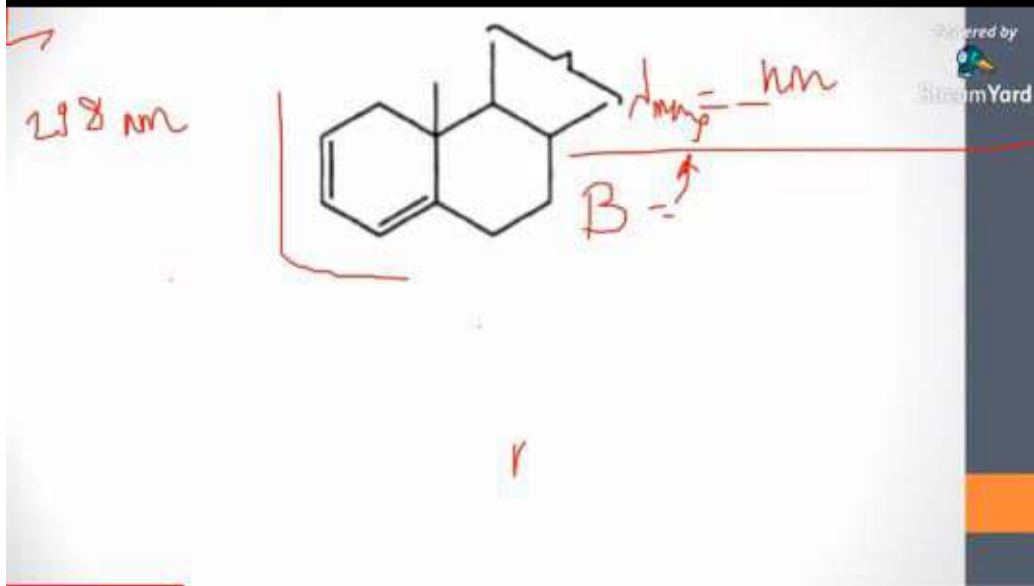
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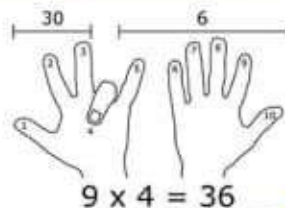


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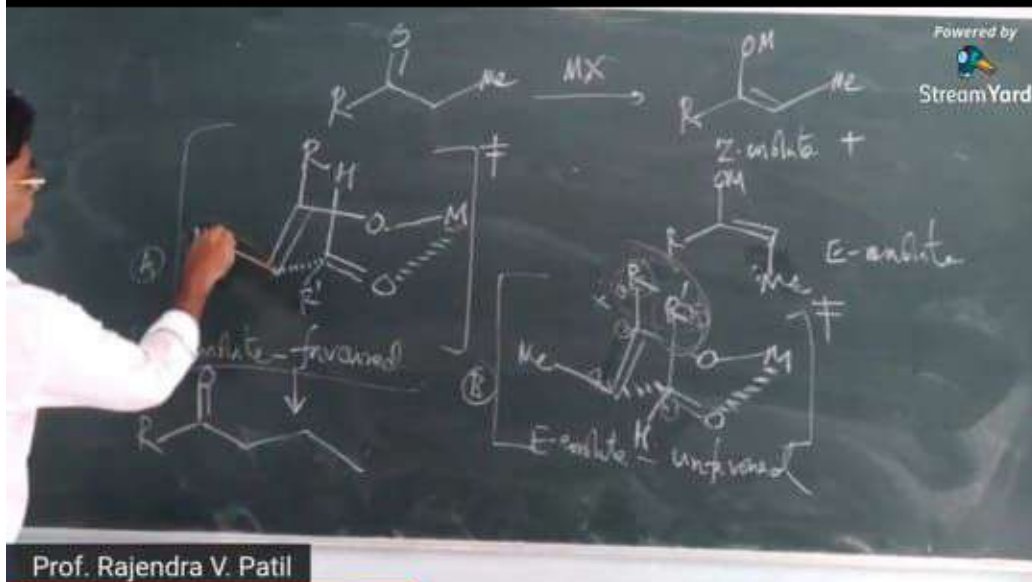
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Prof. Rajendra V. Patil

MSc_CH-352: Asymmetric Synthesis_2



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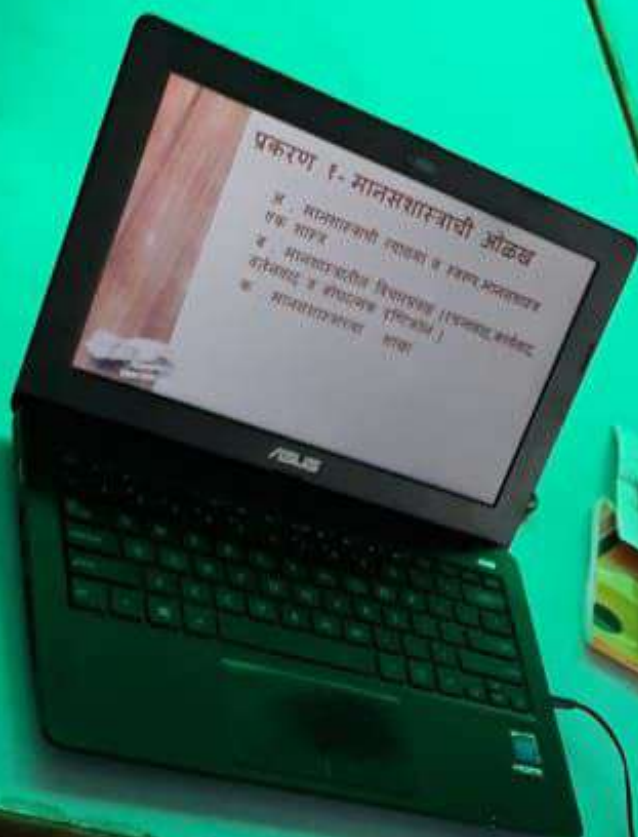
Continuity and Differentiability in One Shot





प्रश्न 1. भारतीय संविधान में 'संघ' का अर्थ क्या है?

- संघ का अर्थ है एक ऐसी संस्था जिसमें दो या दो से अधिक व्यक्ति शामिल हैं।
- संघ का अर्थ है एक ऐसी संस्था जिसमें दो या दो से अधिक व्यक्ति शामिल हैं।
- संघ का अर्थ है एक ऐसी संस्था जिसमें दो या दो से अधिक व्यक्ति शामिल हैं।



7:34 AM

Passive diffusion >

Dr. Varsha Chaudhary

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Dr. V.M. Chaudhary

3:43 / 40:04

Passive and Facilitated diffusion - PowerPoint

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Passive diffusion

Passive diffusion is also known as **simple diffusion**

The transport of molecules across a membrane along the concentration gradient is known as passive transport.

Passive transport does not require an **input** of metabolic **energy**.

The molecule moves from a **higher concentration** to a **lower concentration**.

Passive transport does not require the **carrier** molecule or **integral membrane proteins**.

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Entropy

Entropy is a **quantitative expression** for the randomness or disorder in a system.

Second law introduces the concept of entropy. It is amount of energy unavailable for conversion into work.

The term entropy was introduced by **Clausius** in 1865. It is denoted by Symbol **S**.

Entropy is expressed in terms of entropy unit **EU**, which have dimensions of **Calories/mole/degree**.



Dr. V.M. Chaudhari



Dr. UDDHAV M JADHAV

Starch forms deep blue coloured complex with I_2 but not with I^- ions. Therefore blue colour is absent if solution contains I^- ions.

Solution remains colourless upto equivalence point when Iodine (I_2) is used as titrant against a reducing agent. Because I_2 is reduced to I^- ions. But after equivalence point, when excess of I_2 is added, solution becomes deep blue coloured. It is end point of titration.

But when I^- is used as a reducing agent end point is exactly opposite i.e. blue to colourless.

1.4 Titrations involving Iodine : Iodimetry and Iodometry

In many redox titrations either Iodine is used as a titrant or Iodine is liberated during the titrations. In these titrations starch is used as indicator. These titrations involving Iodine are of two types.

- 1) Iodimetry
- 2) Iodometry

1.4.1 Iodimetry

- Titrations in which Iodine (I_2) is used as an oxidizing agent are called Iodimetric titrations or Iodimetry.
- Iodine is a moderately strong oxidizing agent. Hence it is titrated against strong reducing agents. In these titrations starch is used as indicator.
- In presence of a reducing agent, I_2 is reduced to Iodide ions (I^-). I_2 accepts two electrons from reducing agent to form iodide ions.



Calomel Electrode :- Due the difficulties in the use of standard hydrogen electrode leads to the use of secondary reference electrode. The calomel electrode is widely used because it is easily prepared in laboratory and gives constant emf. There are three types of calomel electrodes.

1. Saturated calomel electrodes
2. 1N calomel electrode
3. 0.1N calomel electrodes

There are different shapes and sizes of calomel electrodes. The commercially available and commonly used is SCE. It consist of a outer glass tube having fine capillary at the bottom of which is plugged with asbestos fiber for connection. The inner glass tube containing mercury and mercurous chloride paste with closed lower end with glass wool plug. A small amount of mercury is placed over the paste and platinum wire is inserted in it for electrical contact. The outer tube is filled with sat. solution of KCl. Some crystals of KCl are placed in outer tube to keep the solution saturated solution of KCl at all temperatures. In case of 1 N and 0.1 N KCl solutions are filled respectively and crystals of KCl are absent. The cell is represented as



The following reaction takes place at the electrode

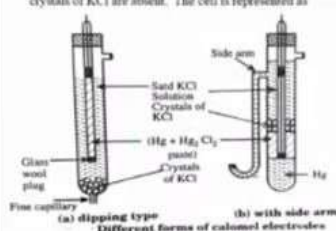


C = saturated or 1 N or 0.1 N

The emf of the electrode is given by the equation,

$$E_{\text{C}} (\text{red}) = E_{\text{C}}^{\ominus} (\text{red}) - \frac{2.303 RT}{F} \log a_{\text{Cl}^-}$$

$$E_{\text{C}}^{\ominus} (\text{red}) = 0.268 \text{ V}$$



(a) dipping type (b) with side arm
Different forms of calomel electrodes

This equation shows that the emf depends on the concentration of chloride ion. The potential of three different electrodes are

Type	Ec (reduction)	Ec(oxidation)	Temp. coefficient
Saturated	0.2415	-0.2415	7.6×10^{-4}
1N	0.2809	-0.2809	2.4×10^{-4}
0.1N	0.3338	-0.3338	7.0×10^{-4}



Dr. UDDHAV M JADHAV



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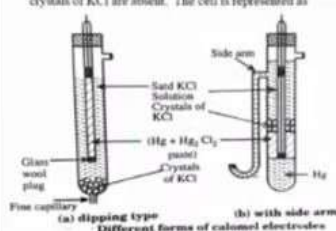


C = saturated or 1 N or 0.1 N

The emf of the electrode is given by the equation,

$$E_{\text{Cl}}^{\text{red}} = E_{\text{Cl}}^{\text{O}} (\text{red}) - \frac{2.303 RT}{F} \log a_{\text{Cl}^-}$$

$$E_{\text{Cl}}^{\text{O}} (\text{red}) = 0.268 \text{ V}$$



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Derivation of Debye-Huckel limiting law :

For a single i^{th} ion showing ideal behavior, the Gibbs energy is given by an expression, $G_i = G_i^0 + kT \ln C_i$, where C_i is its concentration, k is Boltzmann constant. G_i^0 is Gibbs energy of the ion at unit concentration. If the ion is showing deviations from ideality due to ionic atmosphere, the Gibbs energy is then given by

$$G_i = G_i^0 + kT \ln C_i \gamma_i = G_i^0 + kT \ln C_i + kT \ln \gamma_i \quad \dots(8.16)$$

γ_i is the activity coefficient. The third term is the additional term due to the presence of ionic atmosphere. The activity coefficient is calculated as follows :

Imagine that we are able to vary the charge on the ions in solution. Suppose that the charge on all the ions is reduced to zero so that the interactions between the ions disappear and solution becomes ideal. Now imagine that the charge of the ions is increased from zero to their values in the solution. Suppose W is the electrical work done on the system in the charging process. This electrical work being non-PV work is the Gibbs energy change of the process. The work done in charging an isolated ion according to Born model is given by $W = \frac{z_i^2 e^2}{8\pi\epsilon_0\epsilon r} \quad \dots(8.17)$

where r is the radius of ion, e is the electronic charge, z_i is the charge of ion. ϵ_0 and ϵ represent permittivity of vacuum and medium respectively.

Now we require the expression for work of charging the ionic atmosphere which is obtained by replacing r by $1/\kappa$, the radius (thickness) of ionic atmosphere and changing the sign because the net charge of the atmosphere is opposite to that on the ion. Thus,

$$W = -\frac{z_i^2 e^2 \kappa}{8\pi\epsilon_0\epsilon} = \Delta G \quad \dots(8.18)$$

This is the required correction to the Gibbs energy. It is therefore equal to the third term on the right of equation (8.16). We then write

$$kT \ln f_i = -\frac{z_i^2 e^2 \kappa}{8\pi\epsilon_0\epsilon} \quad \text{or} \quad \ln f_i = -\frac{z_i^2 e^2 \kappa}{8\pi kT\epsilon_0\epsilon} \quad \dots(8.19)$$

$$\text{Hence, } \log_{10} f_i = -\frac{z_i^2 e^2 \kappa}{8 \times 2.303 \pi kT\epsilon_0\epsilon} \quad \dots(8.20)$$



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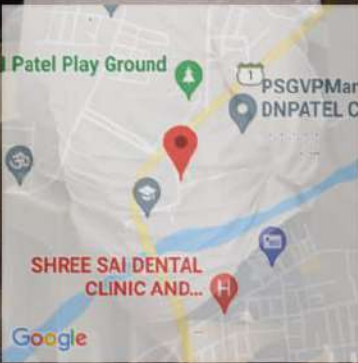


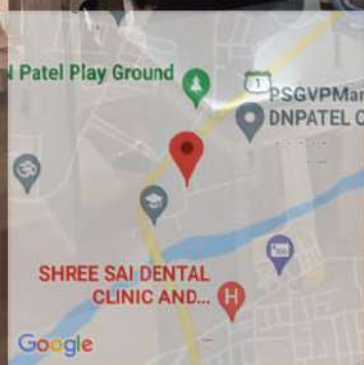
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Nandurbar
Maharashtra
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Jan 6, 2022 12:54:33 PM
314° NW

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Untavad
Nandurbar
Maharashtra
Altitude: 157.6m
Speed: 0.0km/h
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Chap-2 HPLC - LECTURE NO.-1

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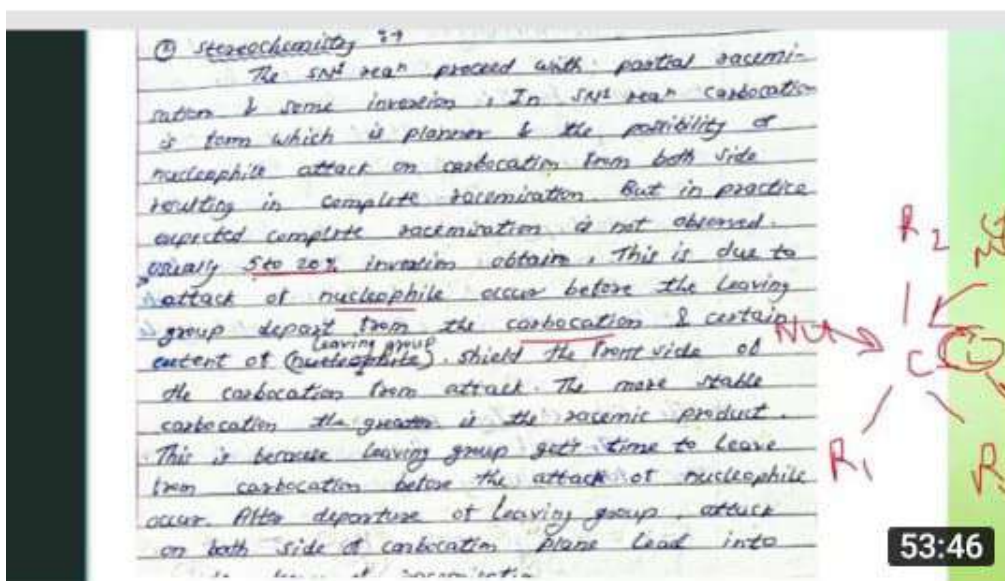
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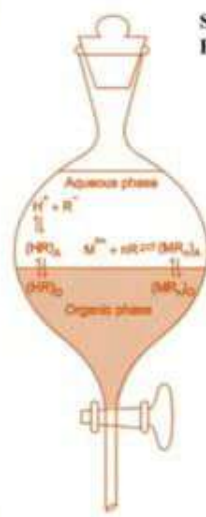


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Nucleophilic Substitution

Mr. Jagdish Chavan · 173 views · Streamed 8



Step - 1 Distribution of chelating agent **HR** between aqueous and organic phase

$$(HR)_o \rightleftharpoons (HR)_a$$

$$k_{D_{HR}} = \frac{[HR]_o}{[HR]_a}$$

$$[HR]_a = \frac{[HR]_o}{k_{D_{HR}}} \leftarrow 1$$

Step-2 Ionization of chelating agent in aqueous phase

$$[HR]_a \rightleftharpoons [H^+]_a + [R^-]_a$$

$$k_a = \frac{[H^+]_a [R^-]_a}{[HR]_a}$$

$$[R^-]_a = \frac{k_a [HR]_a}{[H^+]_a} \leftarrow 2$$

Step-3 the chelating anion $[R^-]_a$ combine with metal ion M^{n+} to form uncharged metal chelate

$$[M^{n+}]_a + [nR^-]_a = [MRn]_a$$

$$k_f = \frac{[MRn]_a}{[M^{n+}]_a [R^-]_a^n}$$

$$[MRn]_a = k_f [M^{n+}]_a [R^-]_a^n \leftarrow 3$$

Step -4 Distribution of metal chelate between aqueous and organic phase

$$[MRn]_a \rightleftharpoons [MRn]_o$$

$$k_{D_{MRn}} = \frac{[MRn]_o}{[MRn]_a}$$

$$[MRn]_o = k_{D_{MRn}} [MRn]_a \leftarrow 4$$

Solvent Extraction - LECTURE NO.-8

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Chemical Equilibrium

Lecture 1

Vishal Tiwari



1:24:13



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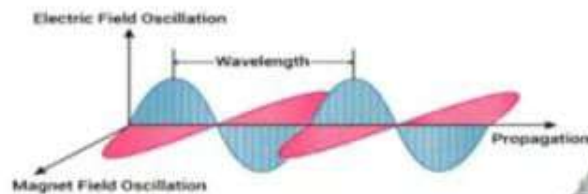
ELECTROMAGNETIC RADIATION AND ITS CHARACTERIZATION

It consists of electric and magnetic fields

It propagates through space or vacuum

It possesses properties of both waves and discrete particles

It is a simple harmonic wave



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Choice of the mobile phase :

The choice of technique depends on the selection of the mobile phase.

- Partition chromatography is used for the separation of highly polar material.
- Adsorption chromatography is used for the separation of non-polar material.
- polarity of different organic compound is as follows :
Amino acid > phenols > H_2O > alcohols > amines > amides > aldehydes > ketones > esters > ethers > nitro compound > halide compounds > hydrocarbon > ionic compounds > oxygenated hydrocarbons.
- The elution is achieved by keeping adsorbent same and increasing the polarity of the solvent. The polarity of different solvent is as follows :
Water > CH_3OH > n-propanol > acetone > ethyl acetate > ethyl ether > chloroform > toluene > cyclohexane > hexane.
- A non-polar mobile phase and polar stationary phase are generally used in reverse phase chromatography.

HPLC



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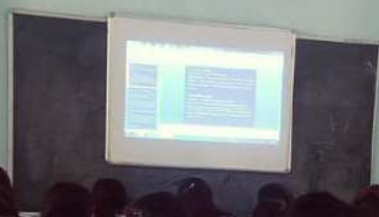
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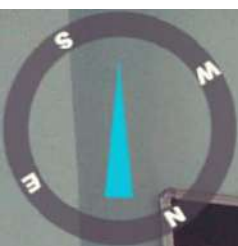
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